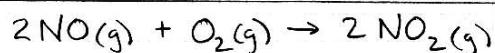
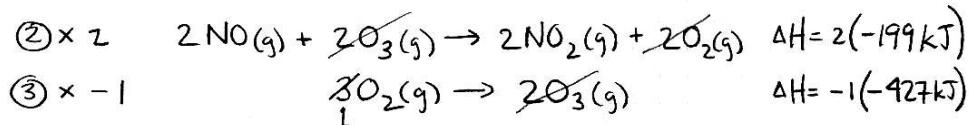
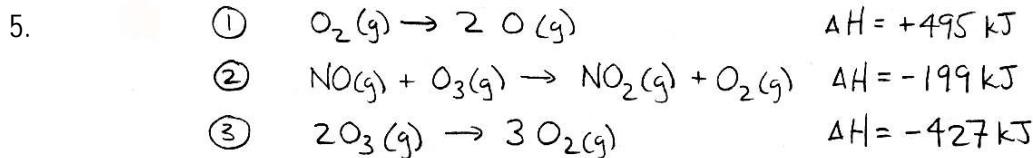
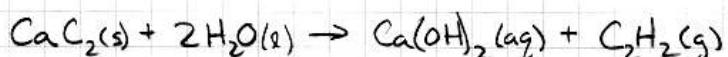
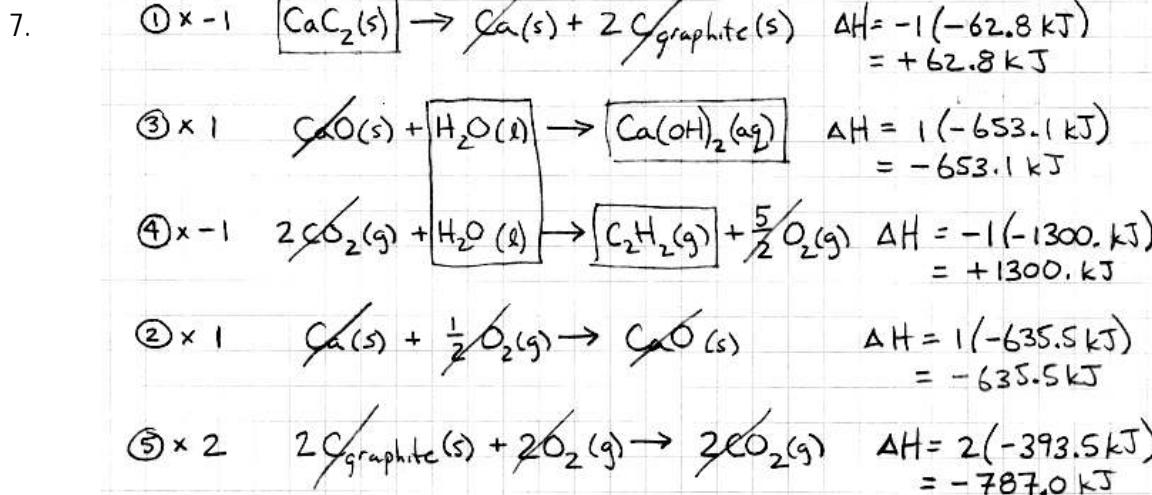


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$$\begin{aligned} \Delta H &= 2(-199 \text{ kJ}) + (-1)(-427 \text{ kJ}) \\ &= +29 \text{ kJ} \end{aligned}$$



$$\begin{aligned} \Delta H &= (+62.8 \text{ kJ}) + (-653.1 \text{ kJ}) + (+1300 \text{ kJ}) + (-635.5 \text{ kJ}) + (-787.0 \text{ kJ}) \\ &= -712.8 \text{ kJ} \end{aligned}$$

*zero decimal places*

so The enthalpy change for the reaction  
is  $-713 \text{ kJ}$ .