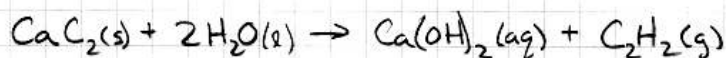
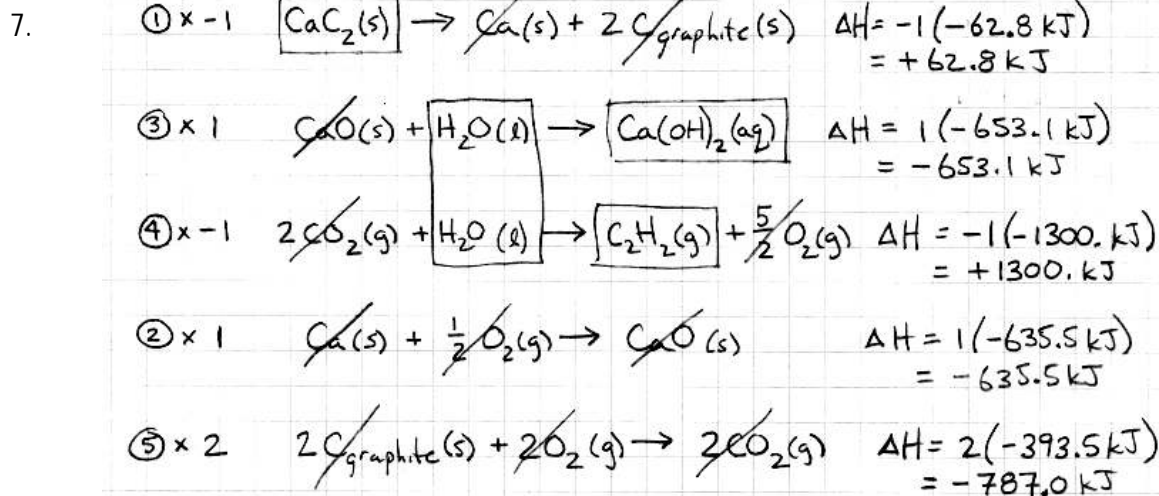


$$\Delta H = 2(-199 \text{ kJ}) + (-1)(-427 \text{ kJ})$$

$$= +29 \text{ kJ}$$



$$\Delta H = (+62.8 \text{ kJ}) + (-653.1 \text{ kJ}) + (+1300. \text{ kJ}) + (-635.5 \text{ kJ}) + (-787.0 \text{ kJ})$$

$$= -712.8 \text{ kJ}$$

zero decimal places

∴ The enthalpy change for the reaction is -713 kJ .